

Name: _____ Per_

Solubility Lab

Prep

Introduction: In this lab you will observe reactions that form solids when solutions are combined together.

Procedure:

1. Obtain a clean, dry well plate and dropper bottles of the following solutions
 - a. Sodium carbonate _____
 - b. Strontium chloride _____
 - c. Ammonium sulfate _____
 - d. Lead (II) nitrate _____
 - e. Potassium iodide _____
2. Write the chemical formulas for the above compounds and get these approved by your instructor before you continue.
3. Combine the solutions in the well plate in the order indicated in the table. Use 5 drops of each solution.
4. Observe the results for each mixture. Record the appearance and make note of which combination resulted in the formation of a solid and which did it.
5. Discard contents of the well plate in the sink and get a clean -up stamp.

Observations:

	Solutions	Precipitate? (yes or No)	Observations (If you wrote yes, describe what the precipitate looks like)
1	Sodium carbonate + strontium chloride		
2	Sodium carbonate + ammonium sulfate		
3	Sodium carbonate + lead (II) nitrate		
4	Sodium carbonate + potassium iodide		
5	Strontium chloride + ammonium sulfate		
6	Strontium chloride + lead (II) nitrate		
7	Strontium chloride + potassium iodide		
8	Ammonium sulfate + lead (II) nitrate		
9	Ammonium sulfate + potassium iodide		
10	Lead (II) nitrate + potassium iodide		

Results: Write the **balanced** equation for each chemical reaction. Include the **state symbols** for each chemical compound. (Hint: You already have the reactant formulas at the top of this paper!)

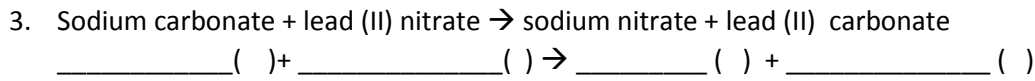
1. Sodium carbonate + strontium chloride \rightarrow sodium chloride + strontium carbonate
_____ () + _____ () \rightarrow _____ () + _____ ()

Which of the products in this equation formed the precipitate? _____

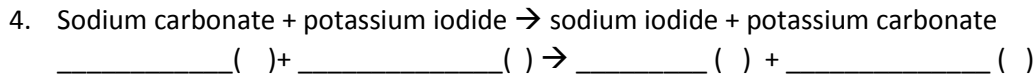
2. Sodium carbonate + ammonium sulfate \rightarrow sodium sulfate + ammonium carbonate
_____ () + _____ () \rightarrow _____ () + _____ ()

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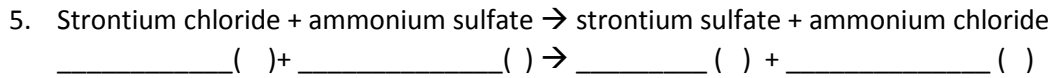
Which of the products in this equation formed the precipitate? _____



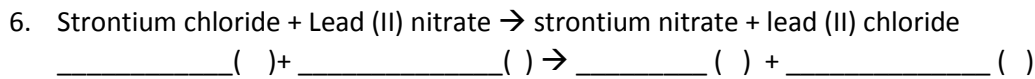
Which of the products in this equation formed the precipitate? _____



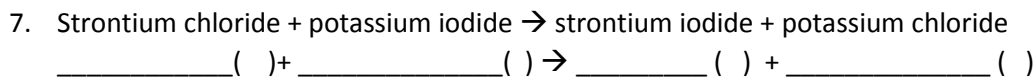
Which of the products in this equation formed the precipitate? _____



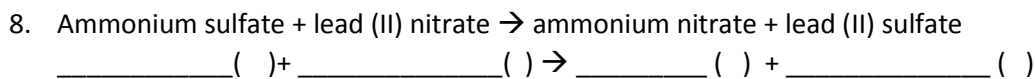
Which of the products in this equation formed the precipitate? _____



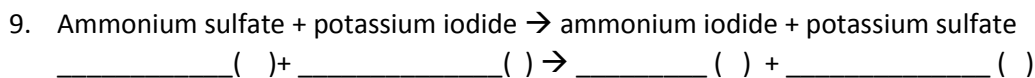
Which of the products in this equation formed the precipitate? _____



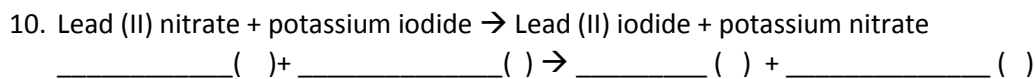
Which of the products in this equation formed the precipitate? _____



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Which of the products in this equation formed the precipitate? _____

Questions:

1. What is the purpose of this lab?
2. What is a precipitate?

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Conclusion: Write one paragraph about your results from the experiment. Write a “mini” discussion. Your mini discussion should include the following

- **Purpose of the lab**
- **Describe the procedure in 1-2 sentences**
- **Your results**
- **Conclusion (Use the sentence frame below)**
- **Sources of error, what you would have done differently.**

I can conclude Claim- what did you learn from the experiment? because Data- the results of your experiment .