

Name: _____ Per _____

Dissociation and Solubility

Practice Sheet #38

1. Write the dissociate equations for the following compounds.

- a. HI
- b. CaBr_2
- c. MgSO_4
- d. $\text{Al}(\text{NO}_3)_3$
- e. CuCl_2
- f. KCN
- g. $\text{Cr}_2(\text{SO}_4)_3$
- h. $(\text{NH}_4)_2\text{SO}_4$
- i. Sodium carbonate
- j. Cobalt (II) chloride

2. Write the dissociation equation for each of the following compounds. Calculate the concentration of each ion in solution.

- a. 2.0 M NaF
- b. 0.50 M BaI_2
- c. 0.40 M K_3PO_4
- d. 0.080 M CrCl_3
- e. 0.10 M $\text{Ni}(\text{NO}_3)_2$
- f. 2.0×10^{-3} M ZnBr_2
- g. 0.050 M $\text{K}_2\text{Cr}_2\text{O}_7$

Name: _____ Per _____

3. Determine whether each of the following compounds is soluble or insoluble in water.

- a. CaF_2
- b. Sodium hydroxide
- c. Iron (II) carbonate
- d. PbBr_2
- e. Ca(OH)_2
- f. Silver chloride
- g. Ammonium phosphate
- h. K_2S
- i. CuI_2
- j. Calcium sulfate

4. Write the formula equation, complete ionic equation, and the net ionic equation for the following reactions.

- a. Silver nitrate and sodium chloride

